Purification of Organic Compounds: Types, Methods & Examples

Simple Crystallisation This is the most common method that we use to purify organic solids. For crystallisation, a suitable solvent is one which dissolves more of the substance at a higher temperature than at room temperature. This solvent is used to dissolve the compound and then allowed to cool slowly. After allowing it to cool slowly, the solid is filtered off and dried.

Recrystallization - Chemistry LibreTexts
Recrystallization, also known as fractional crystallization, is a procedure for purifying an impure compound in a solvent. The method of purification is based on the principle that the solubility of most solids increases with increased temperature.

Crystallization of organic compounds from solution

Laboratory Help! Recristallization of organic compounds
Recrystallization is a widely used purification technique for removing impurities from organic compounds that are solid at room temperature. This method relies on the observation that the solubility of a compound in a solvent generally increases with temperature.

Chapter 12: Recrystallization - Organic Chemistry
Recrystallization is a technique that chemists use to purify solid compounds. It is one of the fundamental procedures each chemist must master in order to perform experiments in the laboratory. Recrystallization is based on the principles of solubility: compounds tend to be more soluble in hot liquids than in cold liquids. This concept is further illustrated in the example of aspirin.

2.1: RECRYSTALLIZATION - Chemistry LibreTexts
Like any purification technique, recrystallization has some limitations. First of all the compound you crystallize should have a high purity and be a crystalline solid. Second, the compound should be insoluble in water. In general the ...